Ch.11-Q1+Q2 Revisin Sheet

Multiple Choice

Identify the choice that best completes the statement or answers the question. 1. The most unreactive group of elements is the ____ c. transition elements a. halogens b. alkali metals d. noble gases 2. According to the law of conservation of matter, if 4.0 g of hydrogen react with chlorine to produce 146 g of hydrogen chloride, how many grams of chlorine reacted? a. 142 g b. 150 g d. 146 g 3. Compare the maximum number of electrons possible in sublevel 3d with the maximum number that could be in sublevel 4d. a. They are the same. c. There are more in 3*d*. b. They are impossible to compare. d. There are more in 4d. 4. Water and hydrogen peroxide are both composed of atoms of hydrogen and oxygen. The differences lie in the arrangement of the atoms. a. composed c. macroscopic b. behavioral d. submicroscopic 5. A certain element is a gas and does not conduct electricity or heat. Which of the following is a possible number of valence electrons for the atoms of this element? a. 2 c. 3 b. 1 d. 6 6. A colorless, odorless gas combines with a magnetic, metallic element. What can you predict about the product? a. It will also be magnetic. b. The compound will be shiny and odorless. c. A gas and a solid produce a liquid. d. It is impossible to predict its specific properties. 7. The inner transition elements are found in the _____ block of the periodic table. a. *d* b. *s* d. *f* 8. Light is released when an electron moves from higher energy levels to a lower energy level. The resulting spectrum is a(n) _____ spectrum. a. excitation c. lower energy d. emission b. absorption 9. Lithium has much less attraction for any valence electrons than does fluorine. Atoms of these two elements would form _____ bonds. a. crystal c. ionic d. molecular b. covalent 10. A physical property of zinc metal is . . a. whether it burns b. how it reacts with nitrogen gas c. its color d. whether it changes when placed into acid 11. An example of a chemical formula is a. d = 13.6 g/Lc. 4.5 g/mL d. H₂SO₄ b. Na

a. periods b. columns d. families 13. Noble gases	Horizontal rows of the periodic table are known as				
13. Noble gases					
0 1 11					
a. form compounds easily					
b. form no compounds that occur naturally in the environment					
c. do not obey the octet rule					
d. form no compounds					
14. In going from left to right in any given row in the periodic table, the size of atoms generally	_ •				
a. changes randomly c. increases					
b. stays the same d. decreases					
15. Which of the following liquids is most volatile?					
a. alcohol c. water					
b. cooking oil d. motor oil					
16. If an atom contains six energy levels, how many sublevels does it contain?					
a. one c. four					
b. six d. two					
17. An example of a pure substance in everyday life is					
a. pond water c. concrete					
b. sugar d. a cola drink					
18. Matter that is large enough to be seen is					
a. macroscopic c. massive					
b. submicroscopic d. a scientific model					
19. The correct way to arrange the three forms of electromagnetic radiation listed below, from highest	t to lowest				
frequency, is					
a. visible > ultraviolet > infrared c. ultraviolet > visible > infrared					
b. infrared > visible > ultraviolet d. infrared > ultraviolet > visible					
20. When reacting with an atom of fluorine, an atom of lithium will lose an electron and become a lit	nium				
a. molecule c. ion					
b. crystal d. compound					
21. Active metals are in the region of the periodic table.					
21. Active metals are in the region of the periodic table. a. p c. s					
21. Active metals are in the region of the periodic table. a. p b. d c. s d. f					
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a. element b. solution d. compound 27. Most transition metals have oxidation state(s). a. no b. only one d. two 28. Transition metals have multiple oxidation states because of the involvement of the electrons in chemical bonding. a. d	 26.	Which of the following materials cannot be broken down into a simpler form?					
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