Q1W3- Qs. Bank- Introduction to periodic table

Multiple Choice

Identify the choice that best completes the statement or answers the question.

 1.	Which of the following elements is a metal?						
	a. Boron	c.	Magnesium				
	b. Nitrogen	d.	Carbon				
 2.	According to periodic table, the physical	and	chemical properties of elements are periodic functions of				
 their atomic weights.							
	a. Dmitri Mendeleev's	c.	Henry Moseley's				
	b. John Newlands'	d.	Lothar Meyer's				
 3.	Which of the following is an example of period	icity	?				
	a. eating breakfast	с.	writing a letter				
	b. hitting a home run	d.	sneezing				
 4.	All of the following elements are metals except						
	a. aluminum	c.	sodium				
	b. chlorine	d.	copper				
 5.	Which element is least likely to be used in semi	icon	ductors?				
	a. silicon	c.	sulfur				
	b. phosphorus	d.	boron				
 6.	is credited with discovering the periodic	law					
	a. Linus Pauling	c.	Dmitri Mendeleev				
	b. Artemis Halogen	d.	J.W. Dobereiner				
 7.	Chlorine, iodine, and make up the haloge	en tr	iad.				
	a. bromine	c.	sodium				
	b. lithium	d.	potassium				
 8.	One of the elements whose existence was predi	cted	by Mendeleev was				
	a. aluminum	c.	potassium				
	b. silicon	d.	germanium				
 9.	The second row of the periodic table includes _		_ elements.				
	a. 2	c.	18				
	b. 8	d.	32				
 10.	is an unreactive element.						
	a. Hydrogen	c.	Helium				
	b. Chlorine	d.	Sodium				
 11.	Which of the following is a transition element?						
	a. gallium	c.	aluminum				
	b. nickel	d.	tellurium				
 12.	All Group 1 elements have						
	a. one valence electron	c.	unpredictable properties				
	b. one energy level	d.	one electron				
 13.	Which of the following events is periodic?						
	a. a basketball game	с.	snowfall				
	b. tides	d.	a single flower blooming				
 14.	Dobereiner's classification system was based on	n gro	oups of elements he called				
	a. tamilies	С.	groups				
	o. perioas	a.	unaus				

 15.	Earliest attempts at classifying elements was based on								
	a. size of atoms	c.	similar properties						
	b. atomic numbers	d.	changing states						
16.	6. The concept of triads suggested that the properties of an element are related to its								
	a. atomic number	c.	periodicity						
	b. atomic mass	d.	melting point						
17.	The blank spaces in Mendeleev's periodic table represented								
 17.	a. liquids	c.	nonexistent elements						
	b. gases	d.	undiscovered elements						
18	Modern periodic law states that properties of el	eme	ints repeat in a regular pattern when the elements are						
 arranged in order of increasing									
	a density	с	atomic number						
	b. atomic mass	d.	periodicity						
10	Horizontal rows of the periodic table are know	n 96	periodicity						
 19.	a groups		 						
	h families	d.	columns						
20	Columns of the noriedie table on known of	u.	columns						
 20.	Columns of the periodic table are known as		similarities						
	a. groups b periods	с. d							
01		u.	10ws						
 21.	Elements in the same group have similar	. .	1						
	a. electron structures	с.	densities						
	b. numbers of electrons	d.	periods						
 22.	At room temperature, most elements are	•							
	a. solid	c.	gas						
	b. líquid	d.	plasma						
 23.	Most elements are								
	a. metals	c.	metalloids						
	b. nonmetals	d.	synthetic						
 24.	Which of the following is not a characteristic o	far	netal?						
	a. lustrous	c.	brittle						
	b. conducts heat	d.	flexible						
 25.	Which groups are considered to be transition el	leme	ents?						
	a. 1 and 2	c.	1, 2, and 18						
	b. 3 through 12	d.	13 through 18						
 26.	Almost all of Earth's atmosphere is made up of		:						
	a. metals	c.	metalloids						
	b. nonmetals	d.	synthetics						
 27.	A certain element is a gas and does not conduc	t ele	ctricity or heat. Which of the following is a possible						
	number of valence electrons for the atoms of the	is e	lement?						
	a. 1	c.	3						
	b. 2	d.	6						
28.	Most semiconductors are .								
	a. metals	c.	metalloids						
	b. nonmetals	d.	synthetics						
29	Which of the following elements is not used to	don	e in an <i>n</i> -type semiconductor?						
 _/.	a. antimony	с.	phosphorus						
	b. arsenic	d.	silicon						

- 30. An element with three valence electrons is used to dope a semiconductor. What type of semiconductor is formed?
 - a. *n* c. *npn* b. *p* d. *pnp*

Completion

Α

Complete each statement.

A- metalloid/s B- alkaline earth C- noble gas

____.

31. The group 2A elements are known as _____ metals.

32. Elements that have physical and chemical properties of both metals and nonmetals are known as

33. A(n) ______ is an element that has properties of both metals and nonmetals.

- 34. An element that appears in Group 18 of the periodic table is called a(n) _____
 - A- periodic law B- period C- semiconductor D- group

35. Family is another name for a(n) ______ of elements.

36. A(n) ______ conducts electricity better than a nonmetal but not as well as a metal.

37. The _______ states that the properties of the elements are a regular function of the elements' atomic numbers.

38. A row of elements in the periodic table is called a(n) ______.

A- transition elements B- periodicity C- lanthanides D- nonmetal

39. The first series of inner transition elements is called the ______.

40. The ______ are the elements that make up Groups 3 through 12 of the periodic table.

41. A(n) ______ is an element in which valence electrons are tightly held.

42. The repeating pattern of the properties of elements from row to row in the periodic table is an example of

A- actinides B- synthetic V- doping

43.	The inner transition e	elements with	atomic number	rs from 90 to	103 make up the	 ·
44.	The conductivity of a	supercondu	ctor can be incre	eased by a pr	ocess known as _	 ·

45. Elements that are not found in nature but are produced artificially are ______ elements.